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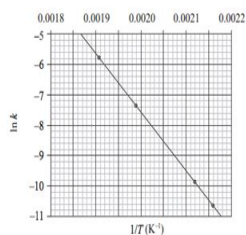


so many fake sites. this is the first one which worked! Many thanks

The conversion of CH_3NC into CH_3CN is an exothermic reaction which can be represented as follows.



This reaction was carried out at different temperatures and a value of the rate constant, k , was obtained for each temperature. A graph of $\ln k$ against $1/T$ is shown below.



- Define the term *activation energy*, E_a . [1]
- Construct the enthalpy level diagram and label the activation energy, E_a , the enthalpy change, ΔH , and the position of the transition state. [3]
- Describe qualitatively the relationship between the rate constant, k , and the temperature, T . [1]
- Calculate the activation energy, E_a , for the reaction, using Table 1 of the Data Booklet. [4]

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